Practice Problems Section 4.1

1) The total concentration of ions in a 0.250 M solution of HCl is \_\_\_\_\_\_\_\_\_\_.

A) essentially zero.

B) 0.125 M

C) 0.250 M

D) 0.500 M

E) 0.750 M

2) A strong electrolyte is one that \_\_\_\_\_\_\_\_\_\_ completely in solution.

A) reacts

B) associates

C) disappears

D) ionizes

3) A weak electrolyte exists predominantly as \_\_\_\_\_\_\_\_\_\_ in solution.

A) atoms

B) ions

C) molecules

D) electrons

E) an isotope

4) Which of the following are strong electrolytes?

 HCl

 HC2H3O2

 NH3

 KCl

A) HCl, KCl

B) HCl, NH3, KCl

C) HCl, HC2H3O2, NH3 , KCl

D) HCl, HC2H3O2, KCl

E) HC2H3O2, KCl

5) Which of the following are weak electrolytes?

 HCl

 HC2H3O2

 NH3

 KCl

A) HCl , KCl

B) HCl , HC2H3O2 , NH3, KCl

C) HC2H3O2, KCl

D) HC2H3O2 , NH3

E) HCl , HC2H3O2 , KCl