Practice Problems Section 4.1

1) The total concentration of ions in a 0.250 M solution of HCl is \_\_\_\_\_\_\_\_\_\_.

A) essentially zero.

B) 0.125 M

C) 0.250 M

D) 0.500 M

E) 0.750 M

2) A strong electrolyte is one that \_\_\_\_\_\_\_\_\_\_ completely in solution.

A) reacts

B) associates

C) disappears

D) ionizes

3) A weak electrolyte exists predominantly as \_\_\_\_\_\_\_\_\_\_ in solution.

A) atoms

B) ions

C) molecules

D) electrons

E) an isotope

4) Which of the following are strong electrolytes?

HCl

HC2H3O2

NH3

KCl

A) HCl, KCl

B) HCl, NH3, KCl

C) HCl, HC2H3O2, NH3 , KCl

D) HCl, HC2H3O2, KCl

E) HC2H3O2, KCl

5) Which of the following are weak electrolytes?

HCl

HC2H3O2

NH3

KCl

A) HCl , KCl

B) HCl , HC2H3O2 , NH3, KCl

C) HC2H3O2, KCl

D) HC2H3O2 , NH3

E) HCl , HC2H3O2 , KCl